Standard Reduction Potentials for Half Reactions
Ionic concentrations, 1M in water at 25°C

Hard to Reduce
Least Reactive

Most Reactive
Easy to Reduce

Easy to oxidize
Most Reactive

Least Reactive
Hard to oxidize

Reduction →				← Oxidation	volts
Li^+	+	e^-	→	Li	-3.00
Cs^+	+	e^-	→	Cs	-2.92
K^+	+	$2e^-$	→	K	-2.92
Rb^+	+	e^-	→	Rb	-2.92
Ba^{2+}	+	$2e^-$	→	Ba	-2.90
Sr^{2+}	+	$2e^-$	→	Sr	-2.89
Ca^{2+}	+	$2e^-$	→	Ca	-2.87
Na^+	+	e^-	→	Na	-2.71
Mg^{2+}	+	$2e^-$	→	Mg	-2.37
Al^{3+}	+	$3e^-$	→	Al	-1.66
Mn^{2+}	+	$2e^-$	→	Mn	-1.18
$2H_2O_{(l)}$	+	$2e^-$	→	$H_{2(g)} + 2OH^-$	-0.83
Zn^{2+}	+	$2e^-$	→	Zn	-0.76
Cr^{3+}	+	$3e^-$	→	Cr	-0.74
Ag_2S	+	$2e^-$	→	$2Ag + S^{2-}$	-0.69
Fe^{2+}	+	$2e^-$	→	Fe	-0.44
Cr^{3+}	+	$2e^-$	→	Cr^{2+}	-0.41
$1/8Se_8 + 2H^+$	+	2e-	→	H_2Se	-0.40
$PbSO_4$	+	$2e^-$	→	$Pb + SO_4^{2-}$	-0.36
Co^{2+}	+	$2e^-$	→	Co	-0.28
Ni^{2+}	+	$2e^-$	→	Ni	-0.25
Sn^{2+}	+	$2e^-$	→	Sn	-0.14
Pb^{2+}	+	$2e^-$	→	Pb	-0.13
Fe^{3+}	+	$3e^-$	→	Fe	-0.04
$2H^+$	+	$2e^-$	→	H_2	-0.00
Cu^{2+}	+	e^-	→	Cu^+	+0.15
Sn^{4+}	+	$2e^-$	→	Sn^{2+}	+0.15
$SO_4^{2-} + 4H^+$	+	$2e^-$	→	$SO_{2(g)} + 2H_2O$	+0.17
Cu^{2+}	+	$2e^-$	→	Cu	+0.34
$I_2(s)$	+	$2e^-$	→	$2I^-$	+0.53
$O_{2(g)} + 2H^+$	+	$2e^-$	→	H_2O_2	+0.68
Fe^{+3}	+	e^-	→	Fe^{2+}	+0.77
Hg^{2+}	+	$2e^-$	→	$Hg_{(l)}$	+0.78
$NO_3^- + 2H^+$	+	e^-	→	$NO_{2(g)} + H_2O$	+0.78
Hg_2^{2+}	+	$2e^-$	→	$2Hg_{(l)}$	+0.79
Ag^+	+	e^-	→	Ag	+0.80
$ClO^- + H_2O$	+	$2e^-$	→	$Cl^- + 2OH^-$	+0.90
$NO_3^- + 4H^+$	+	$3e^-$	→	$NO + 2H_2O$	+0.96
$AuCl_4^-$	+	$3e^-$	→	$Au + 4Cl^-$	+1.00
$Br_{2(l)}$	+	$2e^-$	→	$2Br^-$	+1.06
$O_{2(g)} + 4H^+$	+	$4e^-$	→	$2H_2O$	+1.23
$MnO_2 + 4H^+$	+	$2e^-$	→	$Mn^{2+} + 2H_2O$	+1.28
$Cr_2O_7^{2-} + 14H^+$	+	$6e^-$	→	$2Cr^{3+} + 7H_2O$	+1.33
Cl_2	+	$2e^-$	→	$2Cl^-$	+1.36
Au^{3+}	+	$3e^-$	→	Au	+1.50
$MnO_4^- + 8H^+$	+	$5e^-$	→	$Mn^{2+} + 4H_2O$	+1.52
$4H^+ + SO_4^{2-} + PbO_2$	+	$2e^-$	→	$PbSO_4 + 2H_2O$	+1.68
$H_2O_2 + 2H^+$	+	$2e^-$	→	$2H_2O$	+1.77
$F_{2(g)}$	+	$2e^-$	→	$2F^-$	+2.87